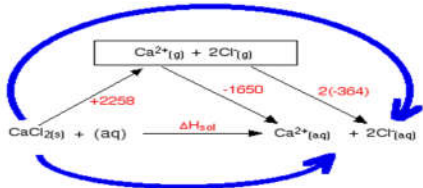
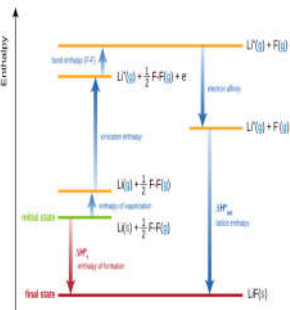


	1	2	3	4	5	6
1	?	entropy	enthalpy of formation	covalent character in ionic compounds	increasing disorder	lattice enthalpy
2	enthalpy of solution	bond enthalpy	$\Delta G = \Delta H - T\Delta S$	electrostatic attraction	?	‘the energy released when 1 mole of gaseous atoms each acquire an electron to form 1 mole of gaseous 1- ions’
3	feasible change	?			plot of ΔG against T	
4	critical temperature	ionic charge and ionic radius	electron affinity	?	always positive	enthalpy of hydration
5	‘the enthalpy change when 1 mole of gaseous atoms is formed from the element in its standard state’			lattice formation	ΔG must be 0 or negative	?
6	?	Gibb's free energy		ionisation energy	‘the enthalpy change when 1 mole of gaseous ions dissolve in sufficient water to give an infinitely dilute solution’	

Energetics