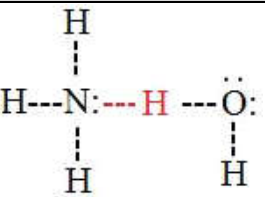
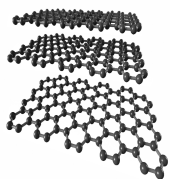
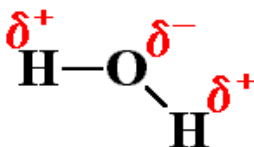
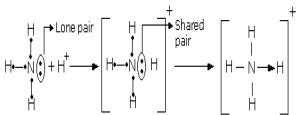
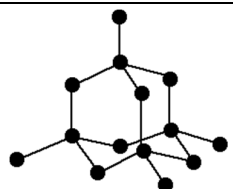


|   | 1                                                                                 | 2                                   | 3                                                                                  | 4                                                                                   | 5                                                                                     | 6                                                                                     |
|---|-----------------------------------------------------------------------------------|-------------------------------------|------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------|---------------------------------------------------------------------------------------|---------------------------------------------------------------------------------------|
| 1 |  | <b>Ionic bonding</b>                | <b>Electrostatic attraction</b>                                                    | <b>Sodium chloride</b>                                                              | <b>Diamond</b>                                                                        | <i>Number of particles in one mole</i>                                                |
| 2 | the power of an atom to attract a pair of electrons in a covalent bond            | <b>Trigonal bipyramidal</b>         |  | <b>Lattice</b>                                                                      | Permanent dipole – dipole forces                                                      | <b>Lone pair-lone pair repulsion</b>                                                  |
| 3 | <b>Iodine</b>                                                                     | <i>Magnesium</i>                    | <b>Bond angle</b>                                                                  |  | <b>Tetrahedral</b>                                                                    |    |
| 4 | <i>Metallic bonds</i>                                                             | <b>Solid less dense than liquid</b> | <b>Lone pairs</b>                                                                  | Van der Waals' forces                                                               | <b>Macromolecule</b>                                                                  | <i>Covalent bonding</i>                                                               |
| 5 | <b>VSEPR theory</b>                                                               | <b>Delocalised electrons</b>        | Co-ordinate (dative) bond                                                          | <b>Graphite</b>                                                                     |  | <b>Pyramidal</b>                                                                      |
| 6 | Electronegativity                                                                 | <b>Permanent dipole</b>             | <i>Ice</i>                                                                         | <b>Polar covalent bond</b>                                                          | <i>Group 2 → 2<sup>+</sup> ion</i>                                                    |  |

## DP Bonding