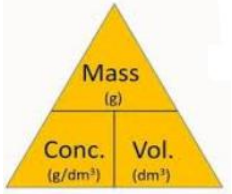
H	1	2	3	4	5	6
1	Number of molecules in 88g of CO_2?	Burning Mg in air results in a product of greater mass	sum of reactant M_r = sum of product M_r	No of electrons in one mole of electrons?	6M	One mole of basketballs
2	Difference between 2CO and CO_2?	M_r	Balanced symbol equations	g/dm^3	Law of conservation of mass	6.02×10^{23}
3		n=C?	24 dm^3	Add 100dm^3 of water to 100dm^3 1M HCl		N
4	Avogadro constant	Balanced symbol equation	Product mass = reactant mass	Mass of 0·1 moles of CH_3CH_2OH?	Mol/dm^3	Relative formula mass
5	$m = \frac{? \times ?}{N_A}$		$NaOH + HCl \rightarrow NaCl + H_2O$ If 25.0 cm^3 of 0.5 mol/dm^3 NaOH reacts with 23.9 cm^3 of HCl, what is the concentration of the acid?		moles	Dilute vs concentrated
6	Loss of gas into the air for example	$NaOH + HCl \rightarrow NaCl + H_2O$ How much sodium chloride can be made by reacting 4g NaOH?		Same volume at same temperature and pressure	State symbols	Molar gas volume at RTP